

**Chemical Bonds:** *holding molecules together*

Type	Description	Example	Diagram
Ionic bond			<p>Sodium ion (+)      Chloride ion (-)</p> <p>11p<sup>+</sup> 11n      17p<sup>+</sup> 18n</p> <p>Attraction between opposite charges</p>
Covalent bond	Non-polar		<p>same charge on both nuclei</p> <p>Electrons spend equal time near each nucleus.</p> <p>(uncharged)</p>
	Polar		<p>Water (H<sub>2</sub>O or H-O-H)</p> <p>(oxygen: slightly negative) (-)</p> <p>larger positive charge</p> <p>Electrons spend more time near larger nucleus.</p> <p>smaller positive charge</p> <p>8p<sup>+</sup> 8n</p> <p>(+) (hydrogens: slightly positive) (+)</p>

Type	Description	Example	Diagram
<i>Hydrogen "bond"*</i>			

\*actually an intermolecular force, not a true bond. But, the name is the name.